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Answers

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Acids, Bases and Salts
Hebden - Unit 4 (page
Page 4/26

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109-182) We will cover
the following topics: 1.

Amphiprotic Substance

2. Amphoteric

Compounds CHEM

0012 Lecture Notes 17

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Hebden - Unit 4 (page

109-182) 1.

Amphiprotic Substance

A substance that can

act either as a

proton acceptor or

a proton donor.

**Acids, Bases and
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Acids, Bases and Salts

SCIENCE AND

TECHNOLOGY Notes

MODULE - 2 Matter in

our Surroundings Acid

Base INTEXT QUESTION

8.1 1. Put the following

substances in acid or

base bottle. (a) Milk of

magnesia (b) gastric

juice in humans (c) soft

drinks (d) lime water

(e) vinegar (f) soap 2.

What will happen if you

add a drop of the

following on a cut ...

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Notes ACIDS, BASES AND SALTS

Unit IV - Acids, Bases
and Salts Notes IV.1 -
The Arrhenius Theory
Of Acids and Bases •

Acid: any substance
which releases H^+
(aq) in water. • Base:
any substance which
releases $OH^-(aq)$ in
water. • Salt: is the
neutralization product
which results when an
acid and a base react:
 $HCl(aq) + NaOH(aq)$
 $NaCl(aq) + H_2O(l)$

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Chemistry 12

Cambridge CIE IGCSE
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4, 6 ranging from 2002
to the present.

Contents include:

Acids, Bases and Salts

Acids and alkalis

Strong ...

**Unit 7: Acids, Bases
& Salts | Teaching
Resources**

Bronsted-Lowry base =

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proton acceptor (H^+ = proton) Bronsted-Lowry acid-base reaction = reaction involving the transfer of a proton
e.g. $HCl + NaOH \rightarrow NaCl + H_2O$
 H^+ donates H^+ \therefore acid OH^- accepts H^+ \therefore base
e.g. $HNO_3 + NH_3 \rightarrow NH_4^+ + NO_3^-$
 H^+ donates H^+ \therefore acid NH_3 accepts H^+ \therefore base

TASK 1 -
Bronsted-Lowry acids & bases

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Chemsheets A2 009

1

A salt is an ionic compound that can be formed by the neutralization reaction of an acid and a base. The following table shows the patterns for the solubility of salts . To help you remember, think of SPANE for soluble salts: Sodium, Potassium, Ammonium, Nitrates and Ethanoates.

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**Acid, Bases, Salts -
IGCSE Chemistry
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01-10), Topic:

Chapter 5.

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Properties Of Acids

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And Bases

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**NCERT Solutions for
class 10 chapter 2
Acids, bases and salt**

Compounds formed from combinations of acids and bases taste salty and are referred to in chemistry as salts. Sweet compounds have characteristics of both acids and bases in the same molecule. We will explore the relationship between molecular structure

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and acids-bases, and consider water solutions of acids and bases.

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demonstrate
distinguish between
acids and bases and
use indicators including
litmus paper, pH paper,
and phenolphthalein

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(indicator in acid-base titrations) to determine their relative pH.

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Class 7 Science
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