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CHAPTER 10: Chemical Quantities  
BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The

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mole was founded by a scientist named Avagadro, and he decided to use the

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Chapter 10 Chemical Quantities Practice Chapter 10: Chemical Quantities. 37 mol of N<sub>2</sub> gas. 1 cup flour + 2 eggs + ½ tsp baking powder 5 pancakes • This relationship can be expressed mathematically. 0 kg × 12 apples/ 1 dozen × 8 seeds/1 apple = 672 seeds Practice Problems Plus How many

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bushels of apples are in 1.

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Chapter 10 - Chemical Quantities. 10.1  
The Mole: A Measurement of Matter 10.2  
Mole - Mass and Mole - Volume  
Relationships 10.3 Percent Composition  
and Chemical Formulas Chapter 10 Test  
Review 10.2 Practice Problems ANS.

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**SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER** (pages 287-296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance.

### **SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)**

Chapter 10 "Chemical Quantities" Vocab. the SI unit representing  $6.02 \times 10^{23}$  representative particles of a substance. the temperature and pressure at which one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

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Worksheet Chemical quantities  
Determine the chemical quantities for  
the following, make sure your answers  
are in correct number of significant  
figures!!! 1 Calculate the number of  
moles for the following quantities: a  $106 \times 10^{23}$  atoms of tungsten b  $3008 \times 10^{23}$   
molecules of Carbon Dioxide 2

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The presentation of information in this chapter assumes that you can already perform the tasks listed below. You can test your readiness to proceed by answering the Review Questions at the end of the chapter. This might also be a good time to read the Chapter Objectives, which precede the Review Questions.  $1.09 \times 10^4 \text{ kg P g P mol P mol P 4O 10 ...}$

### **Chapter 10 Chemical Calculations and equations**

10.2 Mole to Mole & Mole to Volume Relationships \*For all the problems on this page, first find the molar mass of the compound. 1. What is the mass of 9.45 mol of aluminum oxide? mol A ILO + 3 (I A GO q 63,5ù?- 2. What is the



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mass of  $4.52 \times 10^{-3}$  mol of ethylbenzene,  $C_6H_5CH_2CH_3$ ? \*Ethylbenzene is a hydrocarbon that is produced by burning coal.

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Chapter 10 (Chemical Quantities) Test Study Guide The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists.

### **Chapter 10 Chemical Quantities Test Answer Key**

Jun 24 2020 chemical-quantities-chapter-10-test-answer-key 1/5 PDF Drive - Search and download PDF files for free. This is just one of the solutions for you to be successful. Sample Problem. A chemical bond is the attraction between a positive nucleus and negative electrons, or positive and negative ions.  $0.2 \times 10^{23}$  because that's

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how many atoms ...

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