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1.

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For example, a Lewis
acid BF_3 and a Lewis
base OEt_2
(diethylether) form an
adduct $\text{F}_3\text{B}:\text{OEt}_2$.

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The stability increases by the completion of an octet around boron when such an adduct forms. The stability of an adduct is expressed by the equilibrium constant of the reaction. (3.4.10) $A + B \rightleftharpoons K_f A - B$.

3.4: Acid and base - Chemistry

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BRONSTED-LOWRY

ACIDS & BASES

Bronsted-Lowry acid =

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proton donor (H^+ =
proton) Bronsted-Lowry
base = proton acceptor
(H^+ = proton)

Bronsted-Lowry acid-
base reaction =

reaction involving the
transfer of a proton

e.g. $HCl + NaOH \rightarrow$

$NaCl + H_2O$ H^+

donates H^+ \therefore acid OH^-

accepts H^+ \therefore base

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BSAK Chemistry

Define acids, bases

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and neutral solutions,
in terms of hydrogen
ions and hydroxide
ions. ... Click images to

preview the worksheet
for this lesson and the
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Introduction.

Substances can be
classified as acids,
bases or neutral
substances.

Acids, Bases and pH
| Good Science

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Acid base Chemistry of
Aquatic Systems (PDF
163P) This note covers
Basic Concepts of Acid-
Base Chemistry, Acid-
Base Equilibria in Ideal
Solutions, Acid-Base
Equilibria in Real
Solutions, Calculations
involving Acid-base
Equilibria,
Potentiometric
Titrations of Acids &
Bases, CO₂ Equilibria
in Seawater and CO₂ in
the Ocean

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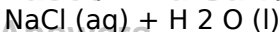
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Chemistry 12 Unit IV –
Acids, Bases and Salts
Notes IV.1 – The
Arrhenius Theory Of
Acids and Bases • Acid:
any substance which
releases H^+ (aq) in
water. • Base: any
substance which
releases OH^- (aq) in
water. • Salt: is the
neutralization product
which results when an

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acid and a base react:



Chemistry 12

According to the Lowry-Bronsted definition, an acid is a proton donor and a base is a proton acceptor. According to the Lewis definition, acids are molecules or ions capable of coordinating with unshared electron pairs, and bases are molecules or ions

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having unshared
electron pairs available
for sharing with acids.

Acids and Bases - Definition, Examples, Properties, Uses ...

R. Janssen, MSEC
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Workbook (Unit 04), P.
8 / 69 32. A 1.0×10^{-4} M
solution has a pH of
10.00. The solute is a...
A. weak acid B. weak
base C. strong acid D.
strong base 33.

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Consider the following
Brønsted-Lowry
equilibrium system... 2
 $\text{HSO}_4^- + \text{H}_2\text{PO}_4^- \rightleftharpoons \text{SO}_4^{2-} + \text{H}_3\text{PO}_4$
4 3 3 4 | What are the
two Brønsted-Lowry
bases in the
equilibrium ...

R. Janssen, MSEC Chemistry 12 Provincial Workbook (Unit 04 ...

Science and
technology makes
good use of acids and
bases. Car batteries

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use a strong acid called sulphuric acid.

Chemical reactions between the acid and lead plates in the battery help make electricity to start the car. They are also used in many household cleaning products, baking soda, and to make fertilizer for crops. Fun Facts. Acids and bases can help neutralize each other. Acids turn litmus paper red, bases turn it blue.

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Strong bases can be slippery and slimy feeling.

Kids science: Acids and Bases

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, Exercise-3 B

Exercise-3 C ,

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decreasing within a
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Questions with
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Chemistry Chapter 2

Important Questions

with Answers Acids,

Bases and Salts. Acids,

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10 Important Questions

Very Short Answer

Type. Question 1. How

is the concentration of

hydronium (H_3O^+)

ions affected when a

solution of an acid ...

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