

## Pogil Phase Changes Answers

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### Pogil Phase Changes Answers

Using the phase diagram o e liquid above, describe any changes in phase present when 1-120 is: 12. kept at OOC while the pressure is increased from 1 atm to 5atm (vertical line) 60b —S 13. Kept at 1.00 atm while the temperature is increased from O oc to 125 oc. (horizontal line)

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Phase Changes Model 2: Temperature of a Substance as Heat is Added Over Time F POGIL 2005, 2006 3/4 Gas boiling D E Liquid B melting C

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Phase Change POGIL Why? Most substances go through a phase change when heated or cooled. Molecules of a substance are held together in either the solid, liquid, or gaseous phase by intermolecular forces. It is necessary to discuss what is occurring at the molecular level in order to explain how an ice cube is melted or how water is boiled.

### 16A-Phase Change POGIL doc - Phase Change POGIL Why Most ...

Most substances will eventually go through a phase change when heated or cooled (sometimes they chemically react instead). Molecules of a substance are held together in either the solidi liquid , or gaseous phase by particle attractions. Though we are very familiar with seeing things like snow melting or water boiling, it is necessary to investigate what is occurring at the particle level in order to truly be able to explain how and why snow melts or water boils.

### Chemistry Crunch #4.1 KEY Phase Changes

Create a cooling curve when given data and identify phase changes, E changes. Model 1 represents molecules in the three phases of matter. Based on what you already know about these phases, complete the table that follows. A word bank has been provided for the last column. D. E. G. F. MODEL 1: PHASE CHANGE PARTICLE DIAGRAM. Word Bank for Column ...

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Phase Changes POGIL 2005, 2006 4/4 Authored by: Erin Graham; Revised by: Josephine Parlagreco, Lizabeth Tumminello Edited by Linda Padwa and David Hanson, Stony Brook University 6. On the molecular level, why is energy added in order to complete an endothermic phase change? (Refer to both Models in your answer.) 7.

### Phase Changes

States of Matter & Phase Changes Why? Most substances go through a phase change when heated or cooled. Molecules of a substance are held together in either the solid, liquid, or gaseous phase by intermolecular forces. It is necessary to discuss what is occurring at the molecular level in order to explain how an ice cube is

### States of Matter & Phase Changes

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### Piersa, Amanda / Matter and Energy

Most substances go thmugh phase change when hentod or cooled. Molecules of a substance are

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held together in either the solid, liquid, or gaseous phase by intermolecular forces. It is necessary to discuss what is occurring at the molecular level in order to explain how an ice cube is melted or how water is boiled.

### Scanned by CamScanner

Phase Changes 6. On the molecular level, why is energy added in order to complete an endothermic phase change? (Refer to both Models in your answer.) 7. On the molecular level, why is energy released to complete an exothermic phase change? (Refer to both Models in your answer.) to 510uJ down, so fly R-ud to 8. Comparing segments

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phase change(s) would occur if I increased the temperature to 600°C? At what temperature(s) would they occur? (NOTE: multiple answers needed for this question) 22) If I had a quantity of substance at a pressure of 2.00 atm and a temperature of -150°C, what phase change(s) would occur if I decreased the pressure to 0.25 atm? At what pressure(s)

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A POGIL Lesson Page 1 of 4 Group Members: Explain and define pressure and how it relates to and applies to the gaseous state. Explain how the quantity, temperature, and volume of the container affect gas pressure. Explain the effect of changes in quantity, pressure, and temperature on the pressure exerted by a gas sample. Model I: Gas Molecules and ...

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(Circle your answer.) i. The temperature decreases as more and more of the liquid is converted into a gas. ii. A liquid boils at a single temperature. iii. The temperature increases as more and more of the liquid is converted into a gas. b. Explain why this observation is reasonable. (Where does the heat energy "go" during a phase change?)

### The Model: Heat and Phase Changes

Using the phase diagram of the liquid above, describe any changes in phase present when H<sub>2</sub>O is: 12. kept at 0°C while the pressure is increased from 1 atm to 5 atm (vertical line) 13. Kept at 1.00 atm while the temperature is increased from 0°C to 125°C. (horizontal line)

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